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| What will we be learning?    Acid and Alkali | Why this? Why now?  Previous units-- Metal and Non-metals, Earth structure and separating mixtures  Next units- Chemical Energy, Types of reactions.  Enquiry Processes  Analyse Patterns, Draw conclusions, Present data, Justify opinions, Collect data, Present data, Plan variables | Key Words:  alkali, acid, neutralisation, salt, hydrochloric acid, sulphuric acid, corrosive, dilute, concentrated, hazard, irritant, indicator, antacid, neutralise, universal indicator, pH scale, chloride, sulphate. |
| What will we learn?  To identify if a substance is an acid or an alkali from its pH number.  Classify solutions as acidic, alkaline or neutral using indicator colours and pH values.  State that in a neutralisation reaction a chemical called a salt and water is formed.  State that in a neutralisation reaction a chemical called a salt and water is formed.  Write a set of instructions to carry out the experiment  Misconceptions in this topic   * Some students may think that antacids are acidic substances and so called because they do not react with acids. * Students will use the term strong/weak and concentrated/dilute to mean the same thing-. A more concentrated solution of a weak acid could, therefore, have a lower pH. An acid’s strength is a measure of how easily the acid dissociates (gives up its hydrogen * Acids and alkalis are dangerous.’ * Students often confuse the concepts of neutralisation and dilution. | |
| What opportunities are there for wider study?  STE(A)M- <https://highcliffe.sharepoint.com/sites/LearnSTEM> | |
| How will I be assessed?  End of topic assessment  Investigation marking  Peer Assessment | |